

# Atomic Structure, Bonding and Types of Substance

**1.32—Substances & their properties**

Structure	Bonding	Relative melting/boiling point (high or low)	Relative solubility in water (high or low)	Relative electrical conductivity as solid (high or low)	Relative electrical conductivity as solution (high or low)
Giant	ionic	High	High	Low	High
Giant	covalent	High	Low	Low* (except graphite/graphene)	n/a (don't dissolve)
Simple	molecular (covalent)	Low	Low	Low	n/a (don't dissolve)
Giant	metallic	High	Low	High	n/a (don't dissolve)

**1.33—Properties of ionic compounds**

Ionic compounds tend to have **high** melting and boiling points.

This is due to the **strong** electrostatic forces of attraction between the oppositely charged **ions**.

These forces require **large** amounts of energy to overcome them.

Ionic compounds can conduct electricity, but only when they are **dissolved** or **molten**, not when they are **solid**.

They can conduct as **solutions / liquids** as the ions are **free** to move and carry a **charge**.

When they are **solid**, the ions are not able to do this.

**1.34—Properties of simple covalent molecules**

Simple covalent molecules tend to have **low** melting and boiling points.

This is due to the **weak** intermolecular forces between the **molecules**.

These forces require **small** amounts of energy to overcome them.

Simple covalent molecules are **unable** to conduct electricity.

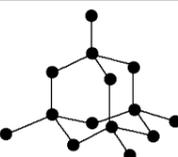
This is because the **molecules** do not have an overall **charge**, and there are no **free electrons**.

**1.35—Diamond and graphite**

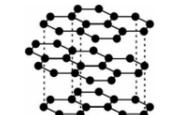
Diamond and graphite are two examples of **giant covalent** substances.

They are two different forms of the same element, **carbon**. *These different forms of the same element are known as **allotropes**.*

**1.36—Structure of diamond & graphite**



**Diamond**



**Graphite**

Describe the structure of diamond. Refer to number & strength of bonds.

- Each C covalently bonds to 4 others; very strong bonds.

Describe the structure of graphite. Refer to the strength and number of bonds, the electrons that don't bond and forces between the layers.

- Each C covalently bonds to 3 others; very strong bonds. Some (unbonded) electrons become delocalised & free to move; weak forces between layers.

**1.37—Properties and uses of diamond & graphite**

Substance	Relative electrical conductivity (high or low)	Relative hardness (hard or soft)
Diamond	Low	Soft
Graphite	High	Hard

Explain why diamond is used in cutting tools. Refer to the table above and any work in section 1.36.

- Very hard, due to the strong covalent bonding of the carbon atoms.**

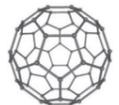
Explain why graphite is used as a lubricant. Refer to the table above and any work in section 1.36.

- Very soft; due to the weak intermolecular forces, the layers can slide easily.**

Explain why graphite is used to make electrodes. Refer to the table above and any work in section 1.36.

- Good electrical conductor, due to delocalised electrons that are free to move & carry a charge.**

**1.38—Properties of other forms of carbon**

Structure	Name	Type of substance	Relative melting/boiling point (high or low)	Relative electrical conductivity as solid (high or low)
	Fullerene	Simple covalent molecules	Low	Low (cannot pass charge between molecules)
	Graphene	Giant covalent	High	High (as with graphite, free moving electrons)

**1.39—Polymers**

Simple (*addition*) polymers, such as poly(ethene), are **large** molecules containing **chains** of **carbon** atoms.

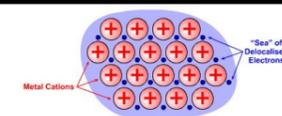
**1.40—Properties of metals**

Explain why metals are malleable, in terms of layers of atoms and forces.

- The layers of positive ions can slide, remain attracted by the negative electrons.**

Explain why metals are good conductors of electricity.

- Free moving (delocalised) electrons which can carry a charge.**



**1.41—Modelling structures**

There are many models, including **dot-and-cross** diagrams, ball-and-stick models and 3d representations. We need multiple models as each individual model has different **limitations**.

**1.42—Metals vs. non-metals**

Metals tend to be **shiny** when cut, have **high** melting points, **high** density and are **good** conductors of electricity. Non-metals tend to be the **opposite**.